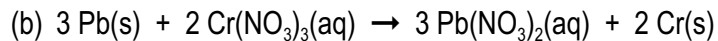
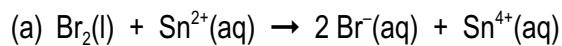
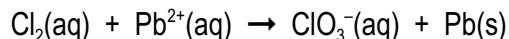


ELECTROCHEMISTRY REVIEW AND PRACTICE 2

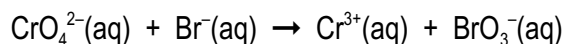
1. For each reaction, write the oxidation and reduction half-reactions, and determine if the reaction is spontaneous.



2. Use the oxidation-number method to balance the following redox reaction occurring in basic solution.



3. Use the half-reaction method to balance the following redox reaction occurring in acidic solution.



4. Using a redox table, predict the reaction that occurs when hydrogen gas is bubbled through an aqueous solution of silver nitrate.

5. Consider the following galvanic cell.



- Draw a complete, neat, and clearly labelled diagram of the cell.
- Label the anode and cathode, and write the half reaction occurring at each.
- Draw a wire connecting the two electrodes and show the direction of electron flow in this wire.
- Show the direction of movement of anions and cations in the solutions.
- Write the balanced net reaction equation for the cell.
- Calculate standard cell potential for this cell.